Hydrogen Storage on Graphene

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National Enterprise for nanoScience and nanoTechnology

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Outline

- Presentation of NEST
- Introduction to Hydrogen Storage
- Epitaxial Graphene
- Hydrogen Storage by Corrugation
- Hydrogen Storage by Functionalization

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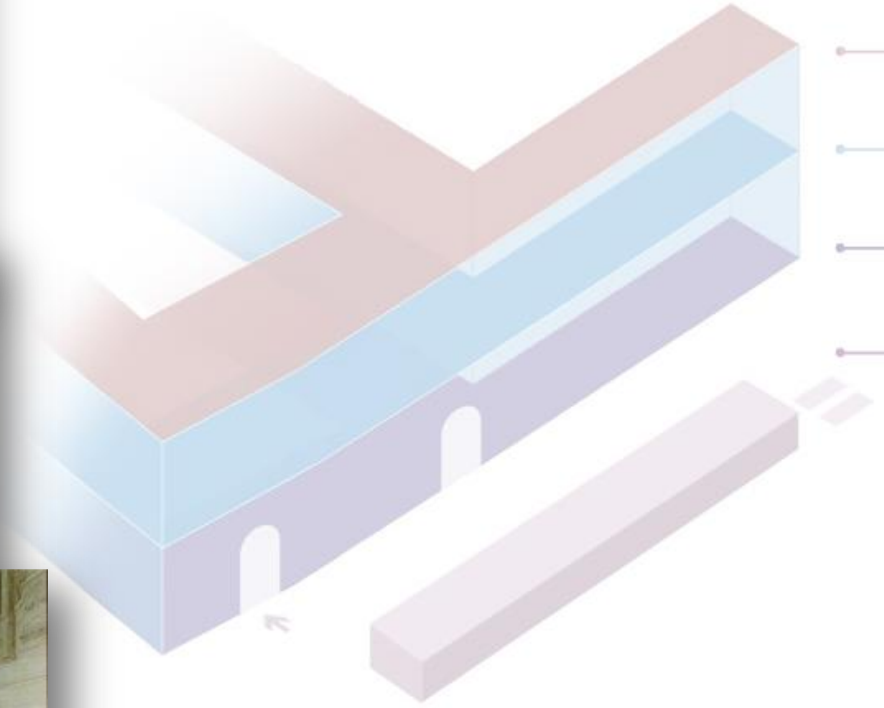

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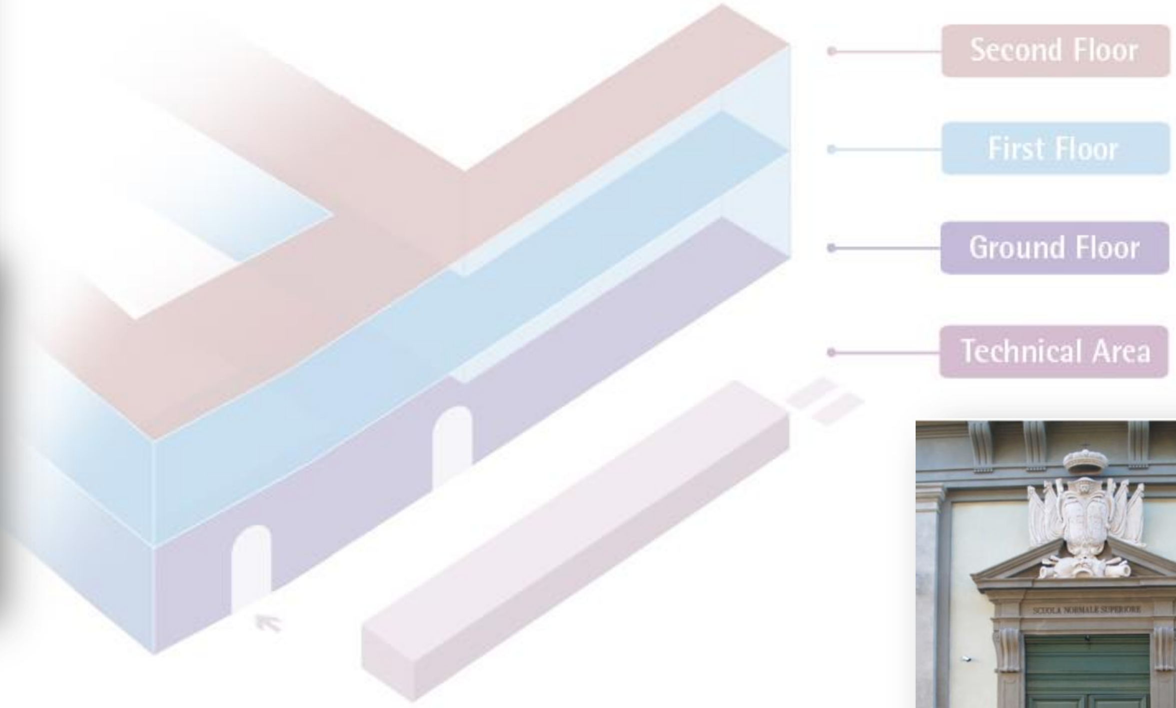

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- **CNR**: Vittorio Pellegrini, Valentina Tozzini, Stefan Heun
- **IIT**: Camilla Coletti, Vincenzo Piazza, Sarah Goler, Torge Mashoff
- **SNS**: Massimo Morandini, Pasquale Pingue, Fabio Beltram

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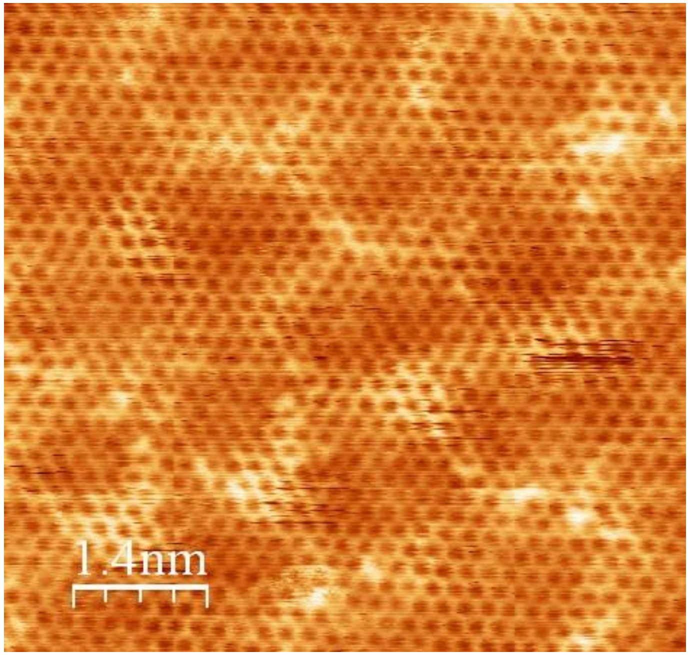
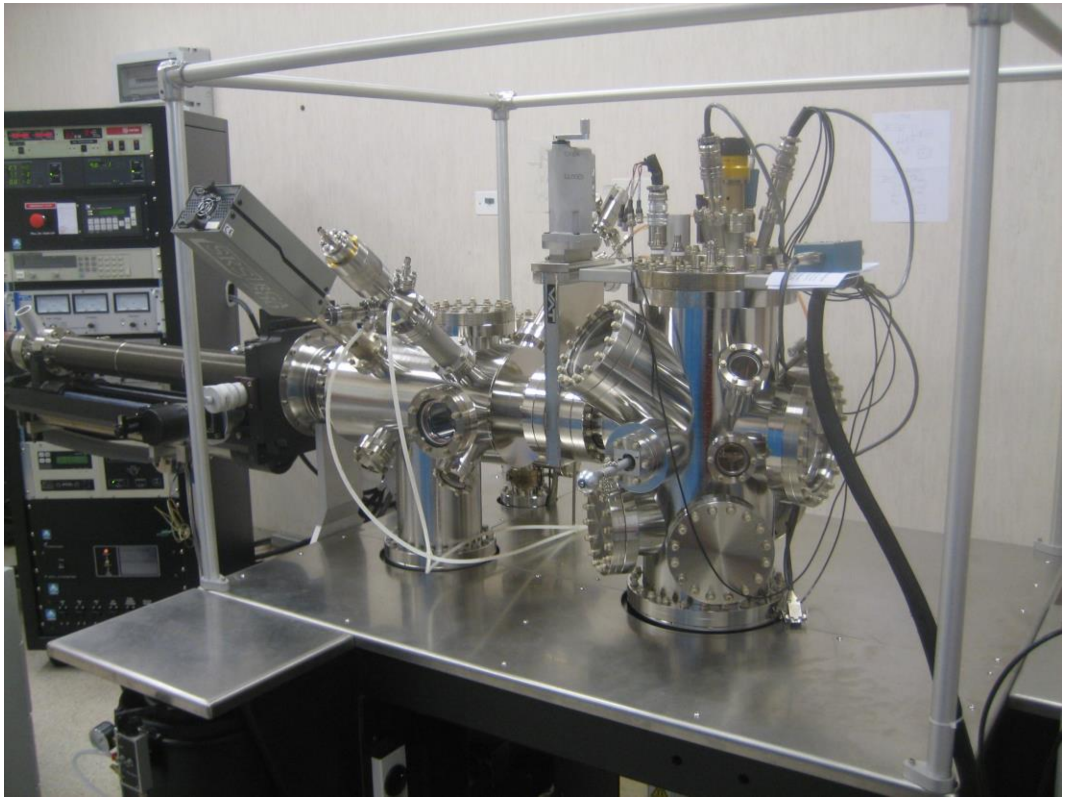
NEST facilities

- UHV VT STM





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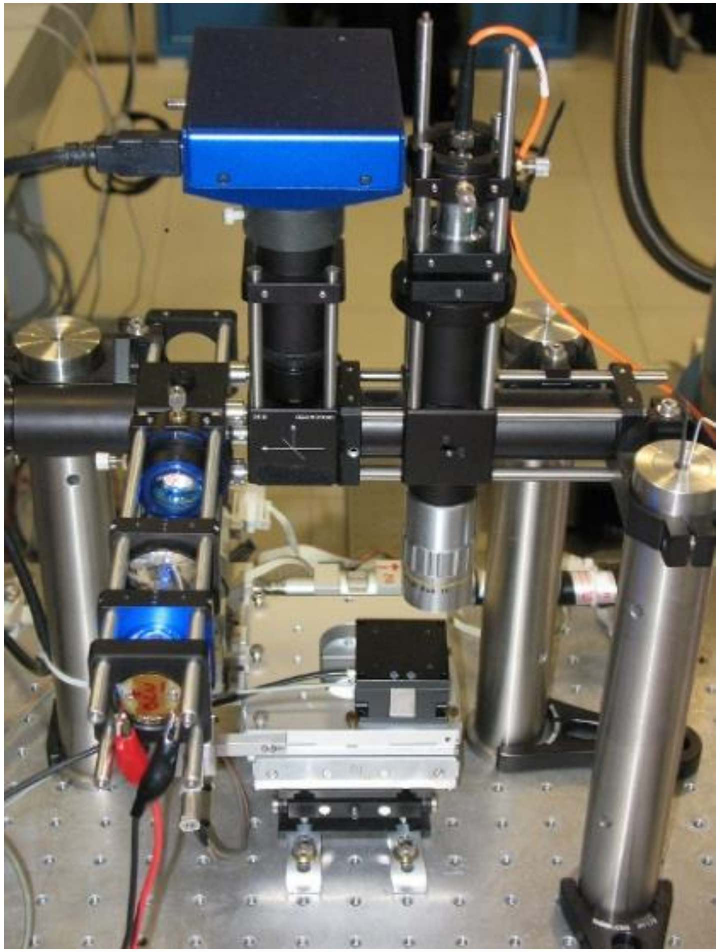
NEST facilities

- UHV VT STM
- Graphene Growth

NEST facilities

- UHV VT STM
- Graphene Growth
- Micro-Raman



NEST facilities

- UHV VT STM
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- Micro-Raman
- AFM

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- AFM
- SEM

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NEST facilities

- UHV VT STM
- Graphene Growth
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- AFM
- SEM
- Theory

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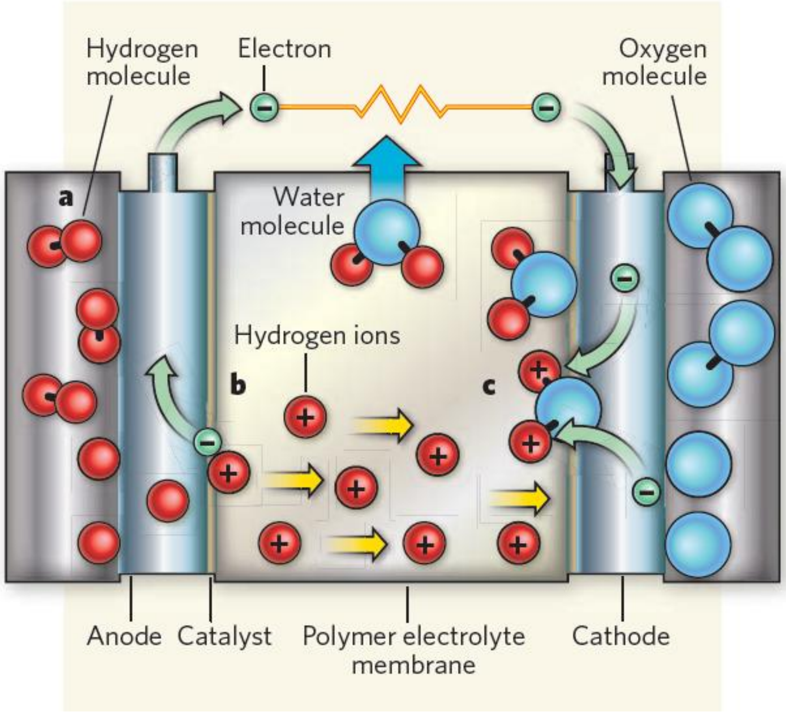
Hydrogen & energy

As a fuel, hydrogen has advantages:

- high energy-to-mass ratio
- $H_2 + 1/2\ O_2 \rightarrow H_2O$ $\Delta H = -2.96eV$
- Non-toxic and "clean" (product = water)
- renewable

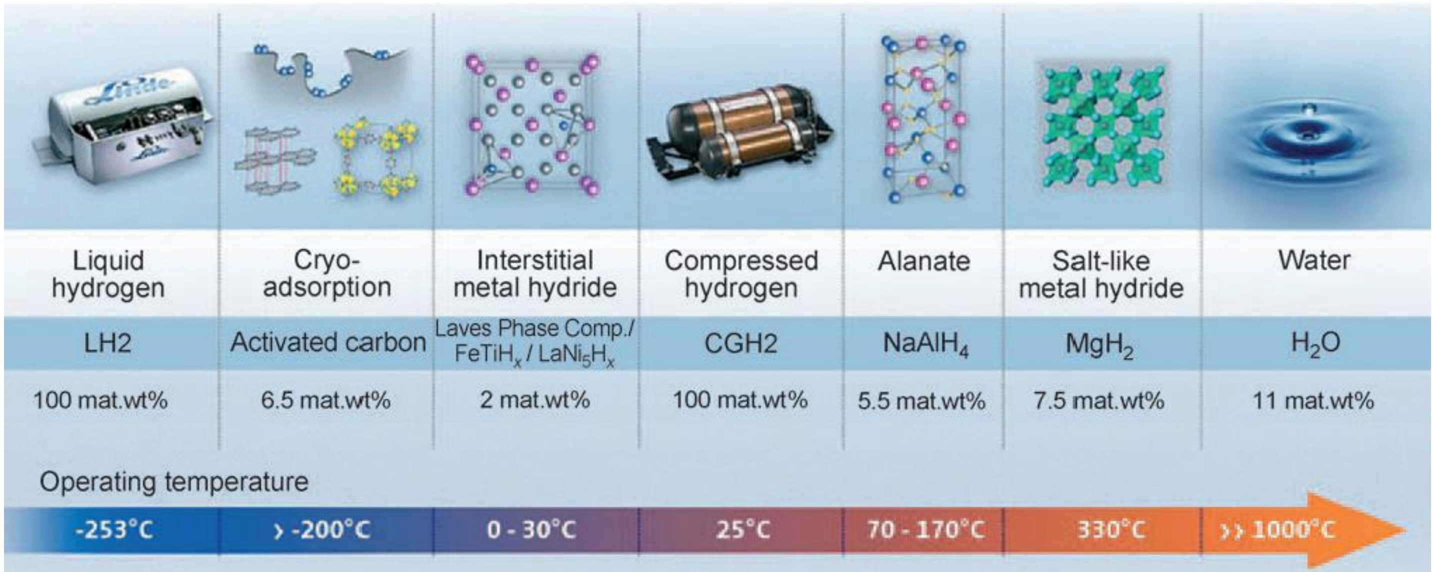
However, hydrogen is NOT an energy source: it must be produced e.g. by electrolysis, needing +2.96 eV, with zero balance with respect to energy production.



Hydrogen fuel cell

Hydrogen-fuelled vehicles

Hydrogen Storage



Targets for **transport applications** not reached yet:

$\rho_m > 5.5$ wt%

$\rho_V > 50$ kg H_2 /m^3

$P_{eq} \approx 1$bar at T< 100°C

Compressed H_2:
High pressure and heavy container to support such pressure

Solid State:
Physisorption
Chemisorption

Liquid H_2:
Liquefation needs energy and consumes more than 20% of the recoverable energy

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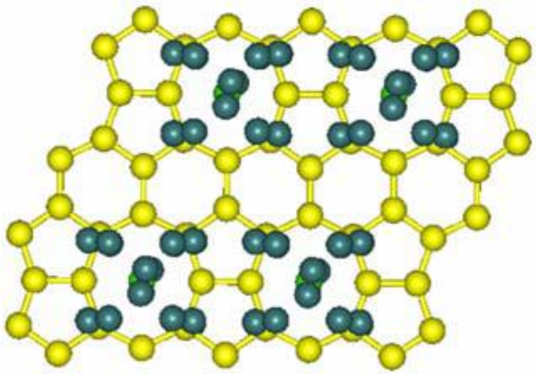
Graphene for hydrogen storage

- Graphene is lightweight, inexpensive, robust, chemically stable
- Large surface area (~ 2600 m^2/g)
- Functionalized graphene has been predicted to adsorb up to 9 wt% of hydrogen

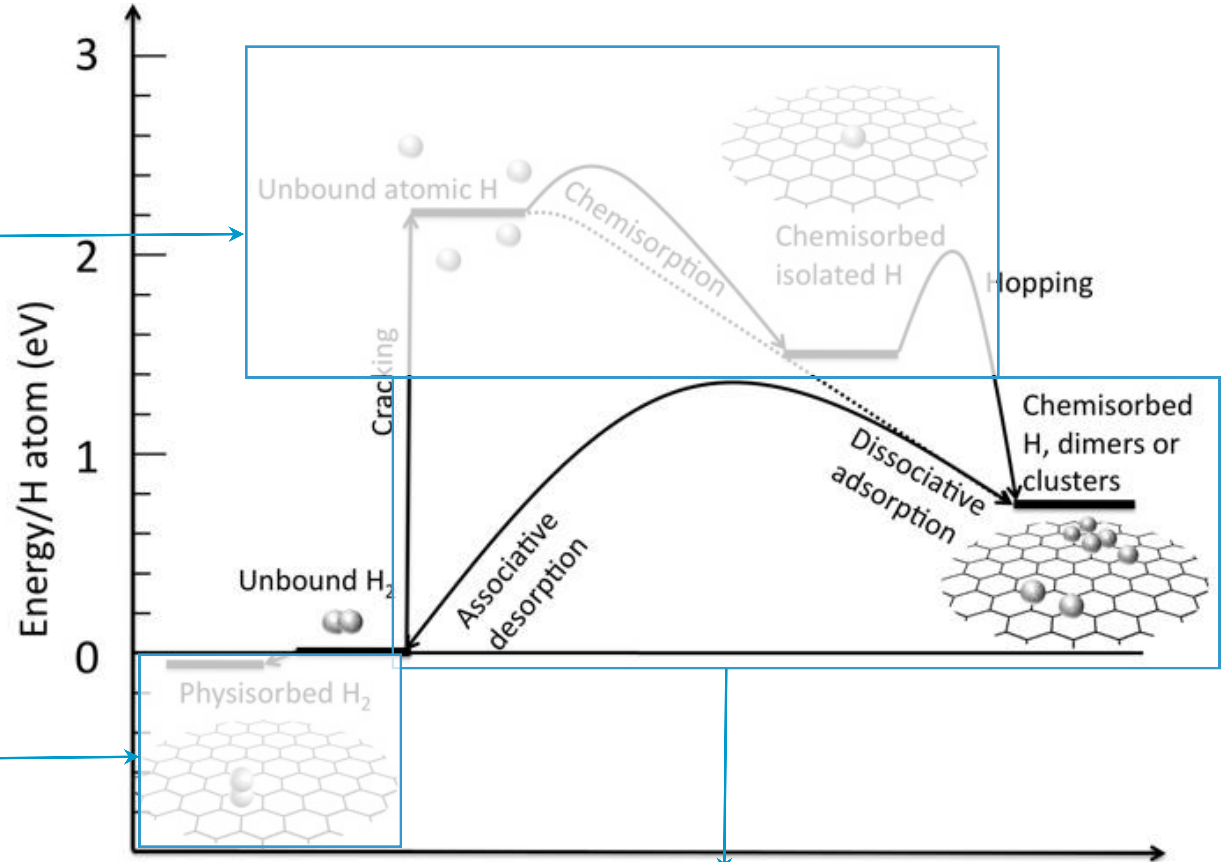


Yang et al., PRB 79 (2009) 075431

H storage in graphene

❖ Atomic hydrogen chemisorption has a small or negligible chemisorption barrier ⇒ feasible but H_2 must be cracked

❖ Physisorption weakly bounds hydrogen ⇒ acceptable storage densities only at low temperatures and/or high pressure



❖ Molecular hydrogen chemi(de)sorption has high barrier (theoretical estimate ~eV) ⇒ chemisorbed H is stable for transportation etc, but catalytic mechanisms are necessary in the loading-release phases

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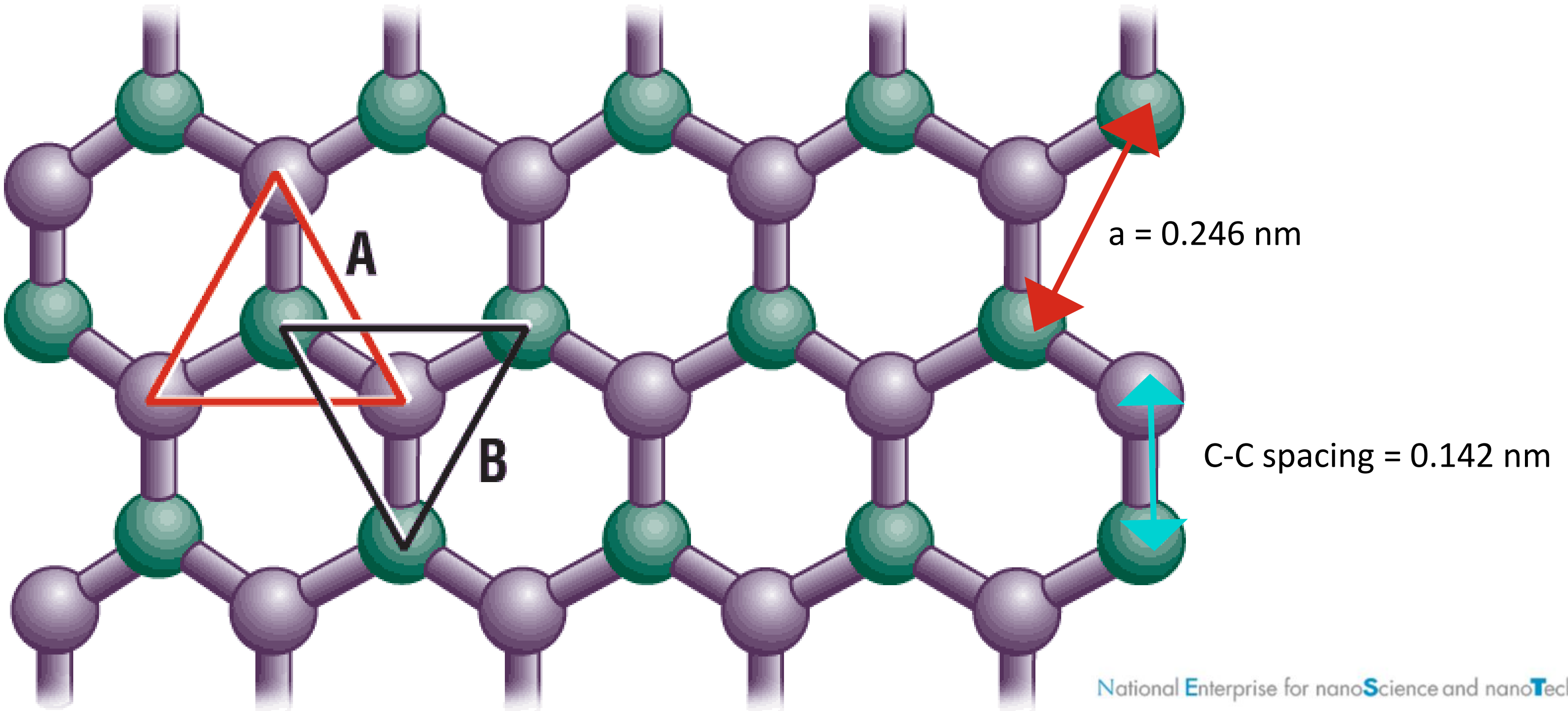
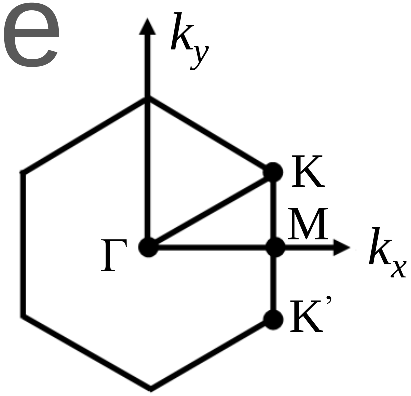
What is Graphene?

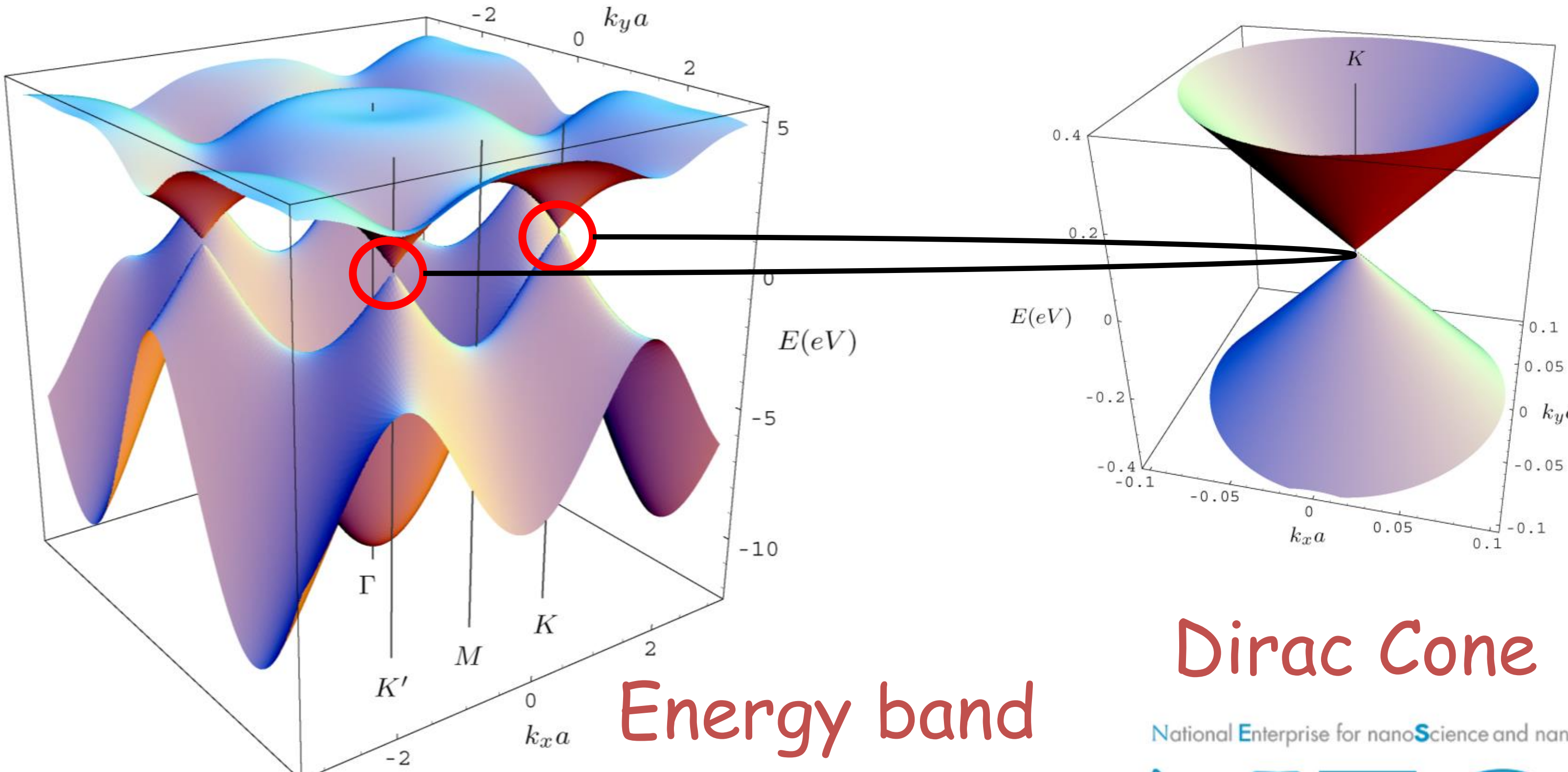
A SINGLE layer of carbon atoms!

The atoms are arranged in a honeycomb lattice composed of two intertwined equivalent sublattices.



The Brillouin Zone



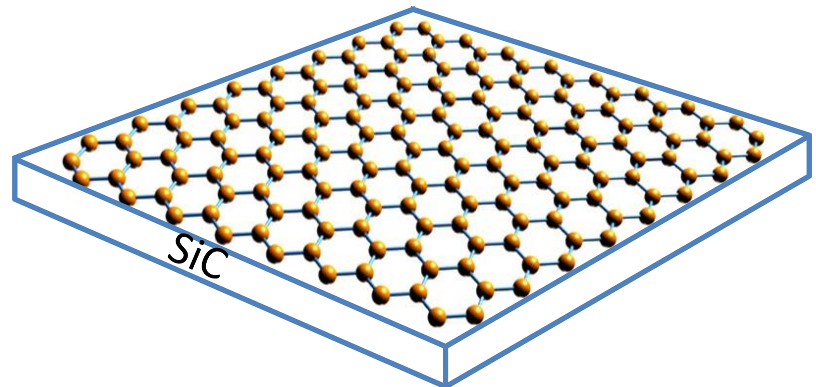


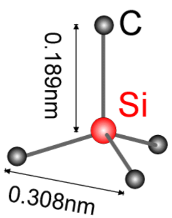
Energy band structure

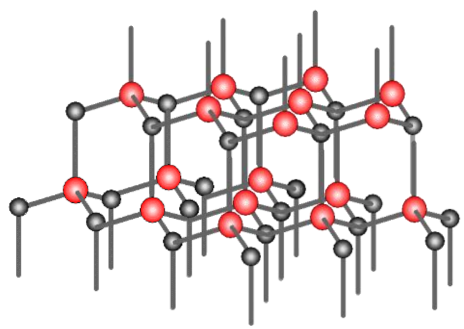
Dirac Cone

Graphene growth on SiC

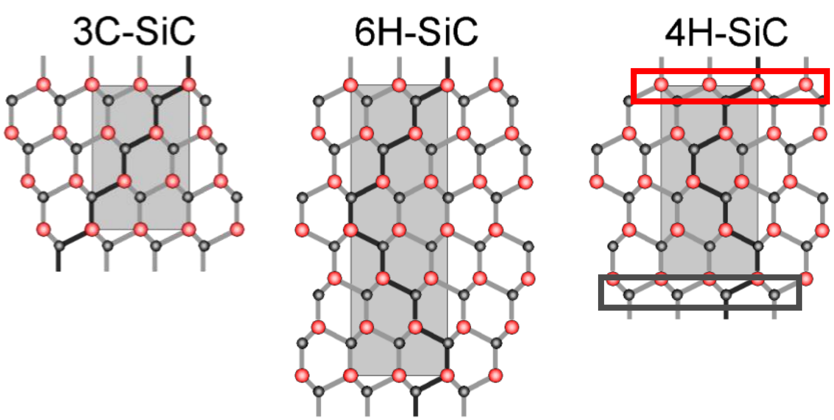
Graphene or thin graphite films form on SiC surfaces upon annealing at high temperatures as a result of SiC decomposition.







Bilayer of Si-C tetrahedra



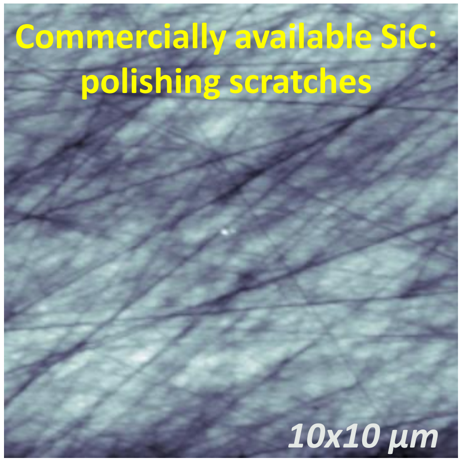
Si(0001) face →

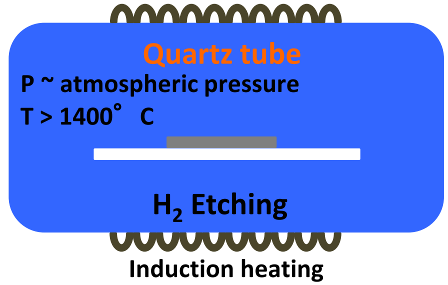
Graphene:
Ordered stacking
Good thickness control

$C(000\bar{1})$ face →

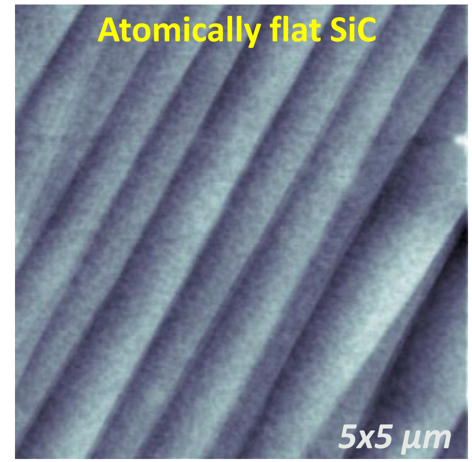
Graphene:
Rotational disorder
Poor thickness control

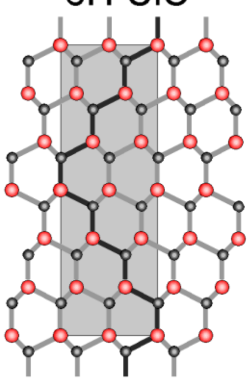
Graphene growth on SiC(0001)

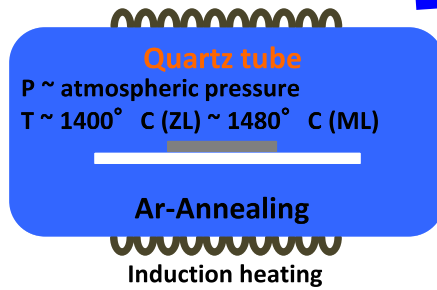




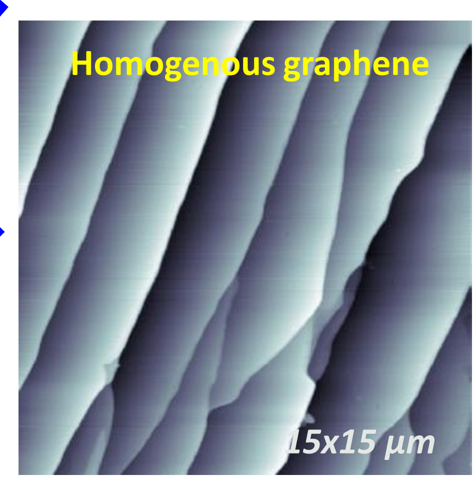
C. Coletti et al., *Appl. Phys. Lett.* 91, 061914 (2007)



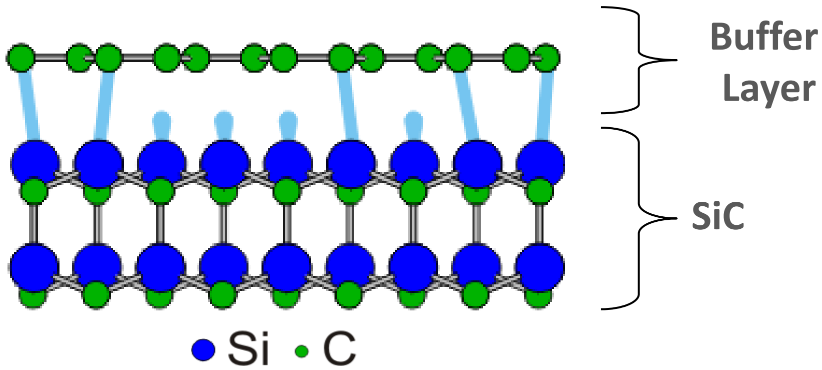


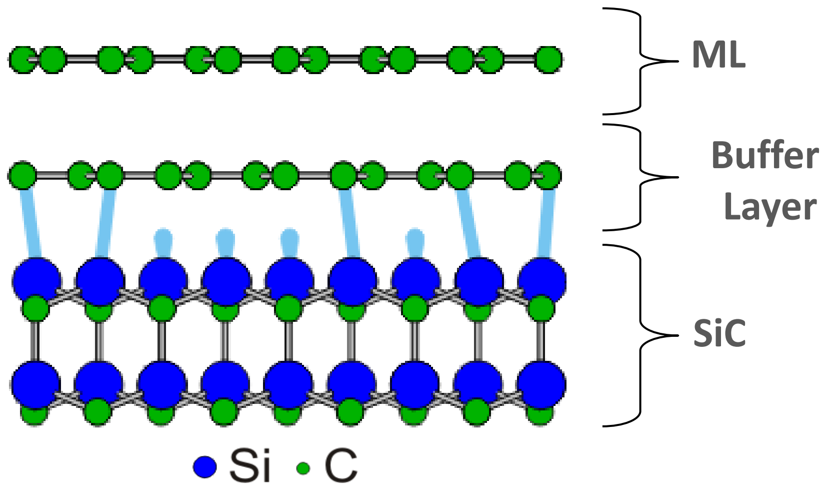


K.Emtsev et al., Nature Mater. 8, 203 (2009)

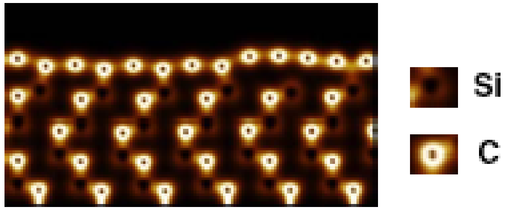


Graphene growth on SiC(0001)



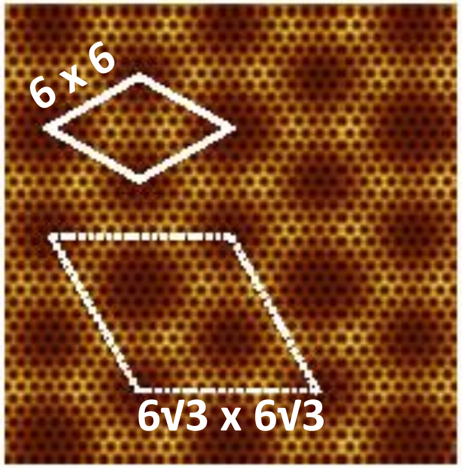


Buffer Layer
Topologically identical atomic carbon structure as graphene. Does not have the electronic band structure of graphene due to periodic sp^3 C-Si bonds.



F. Varchon, et al., PRB 77, 235412 (2008).

Theoretical Calculations



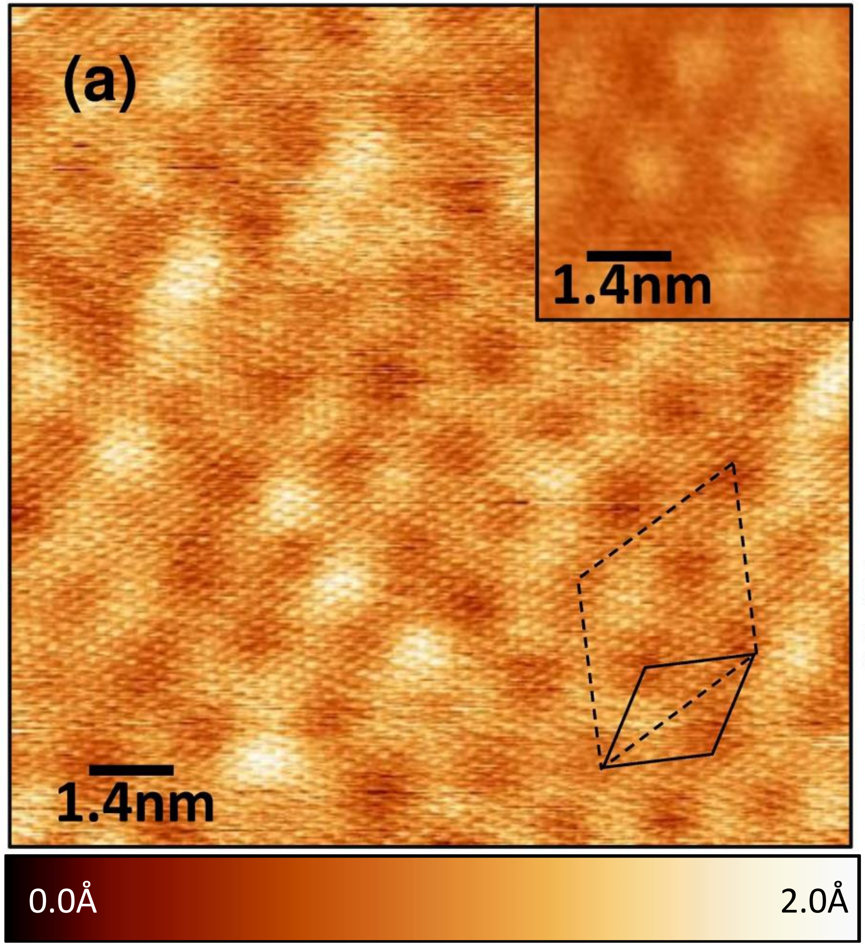
F. Varchon, et al., PRB 77, 235412 (2008).

Superstructure of both the zerolayer and monolayer graphene on the Si face from the periodic interaction with the substrate.

Buffer Layer

Bias: +1.7V
Current: 0.3nA

$6\sqrt{3} \times 6\sqrt{3}$ - - - - -
quasi-(6x6) ——



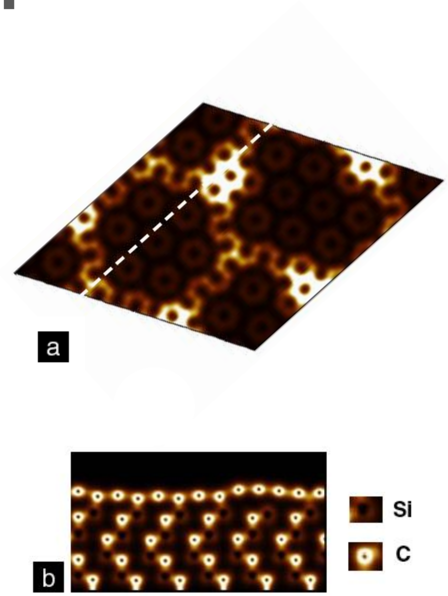


FIG. 2. (Color online) Total charge density of the buffer layer on SiC(001). (a) total charge density in the 6*R*3-SiC unit cell. (b) cross section of the total charge density along the line defined in (a). The black dots that appear when the cross section goes through the middle of an atom are due to the use of pseudopotentials (no core electrons).

F. Varchon, et al., PRB 77, 235412 (2008).

S. Goler *et al.*: Carbon **51**, 249 (2013).

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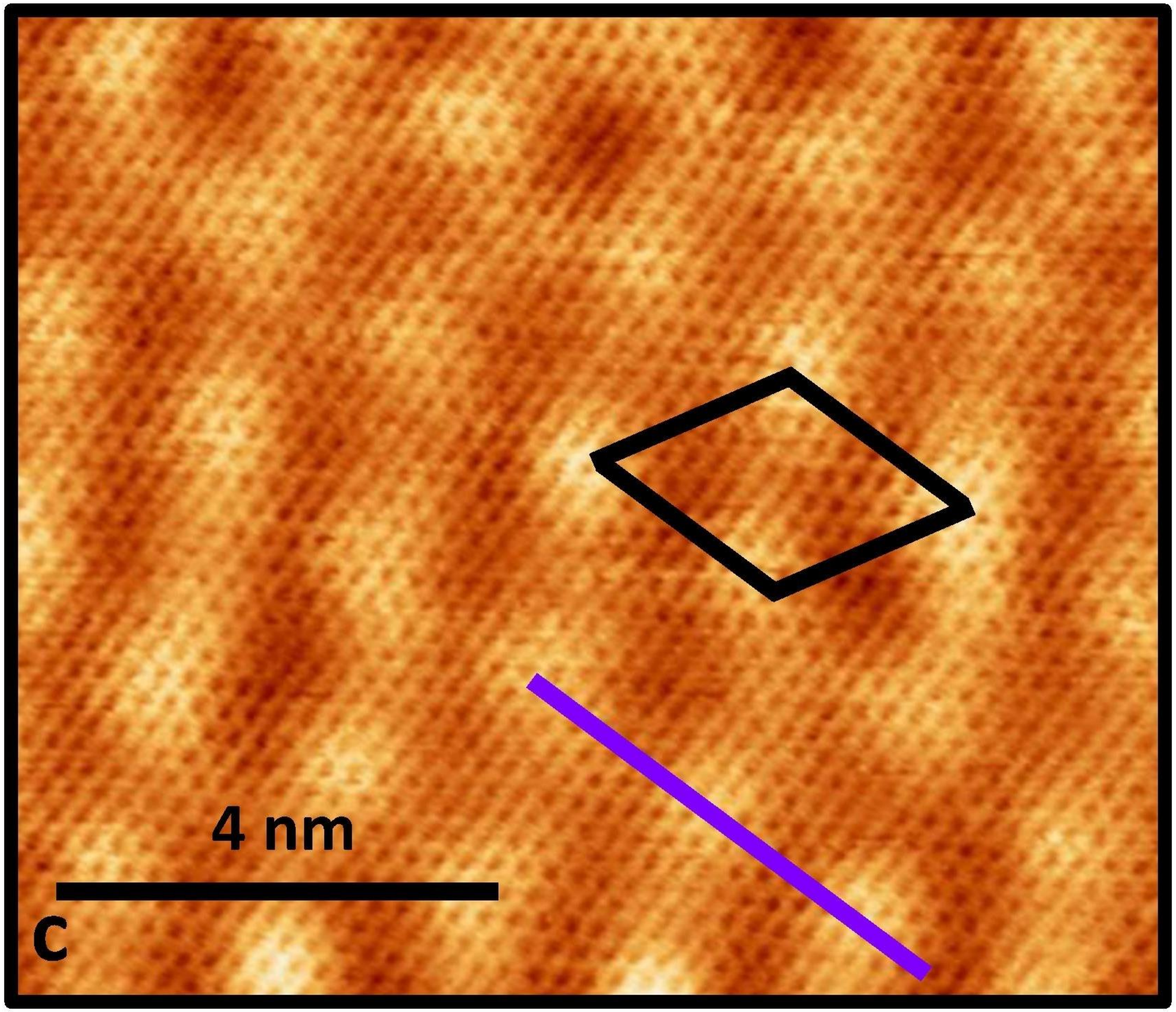
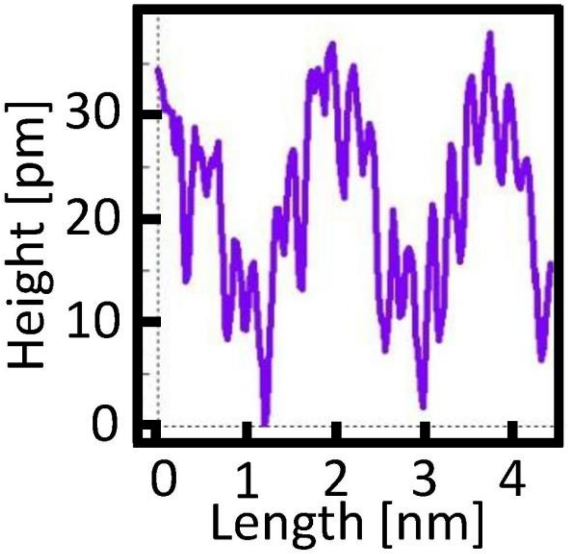
Monolayer Graphene



STM



S. Goler *et al.*

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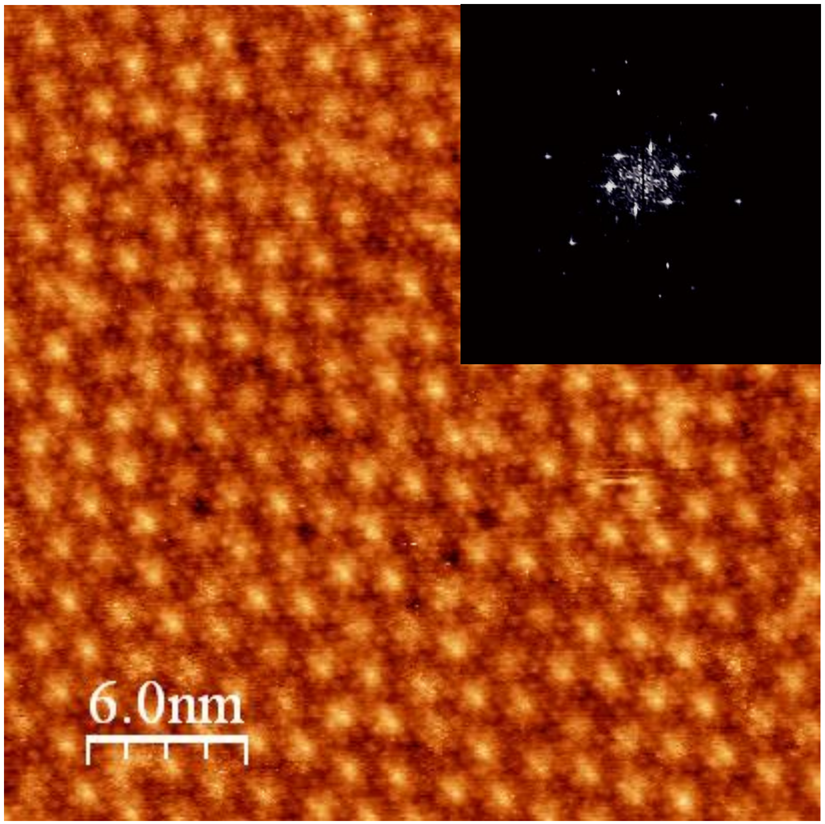
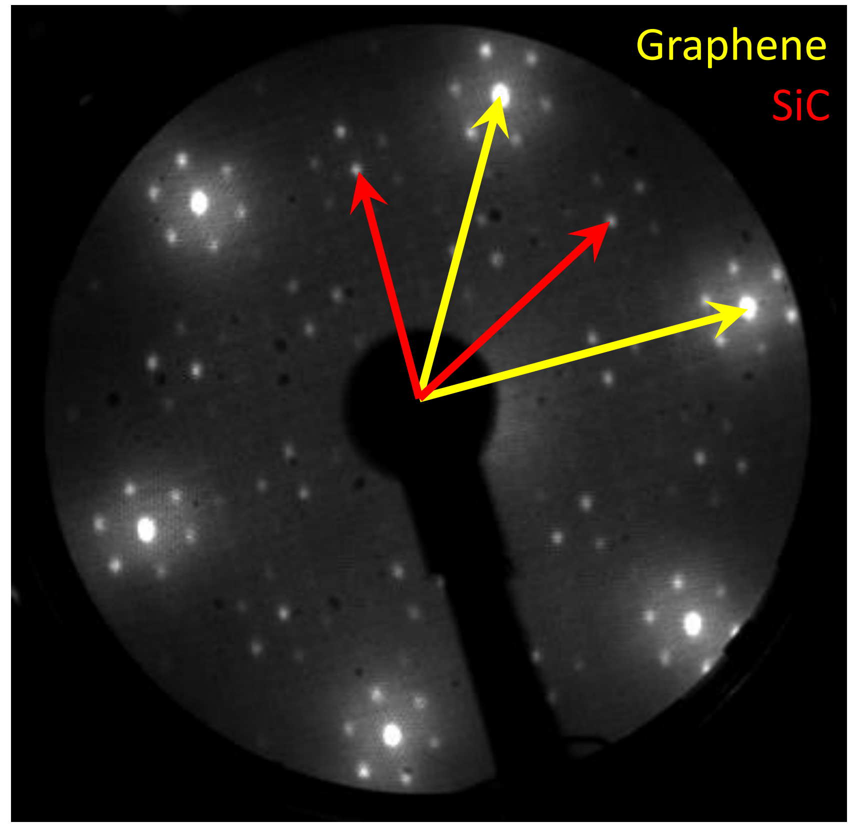
NEST

6√3x6√3-Superstructure

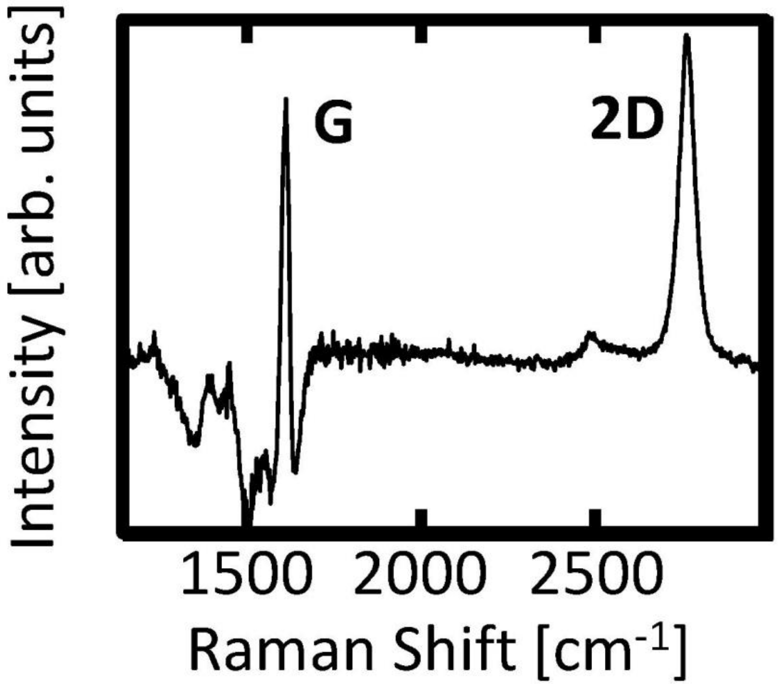


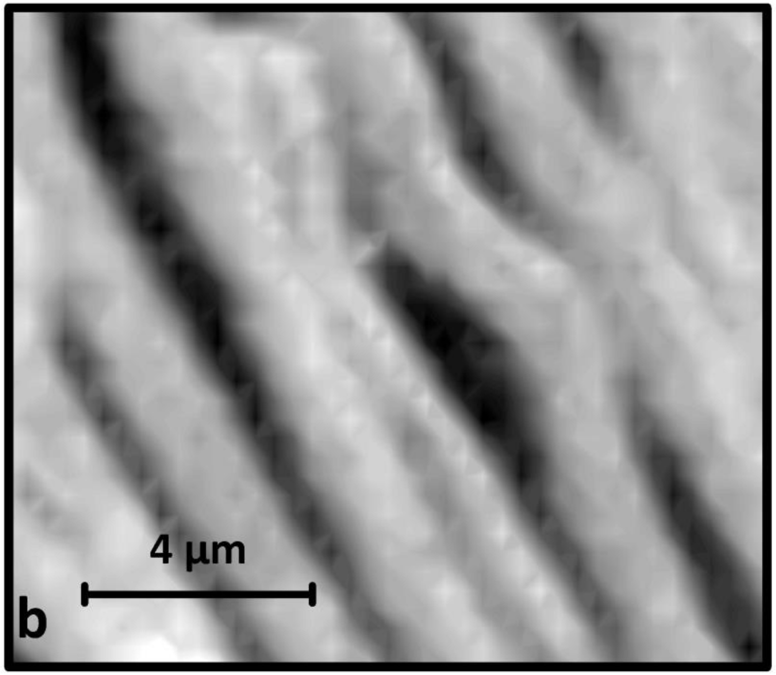
30 nm, 1V, 100 pA



E= 75 eV

ML: Micro-Raman





Spectrum from 12um x 12um area
SiC background subtracted

Integrated intensity of 2D peak
Bright = ML graphene

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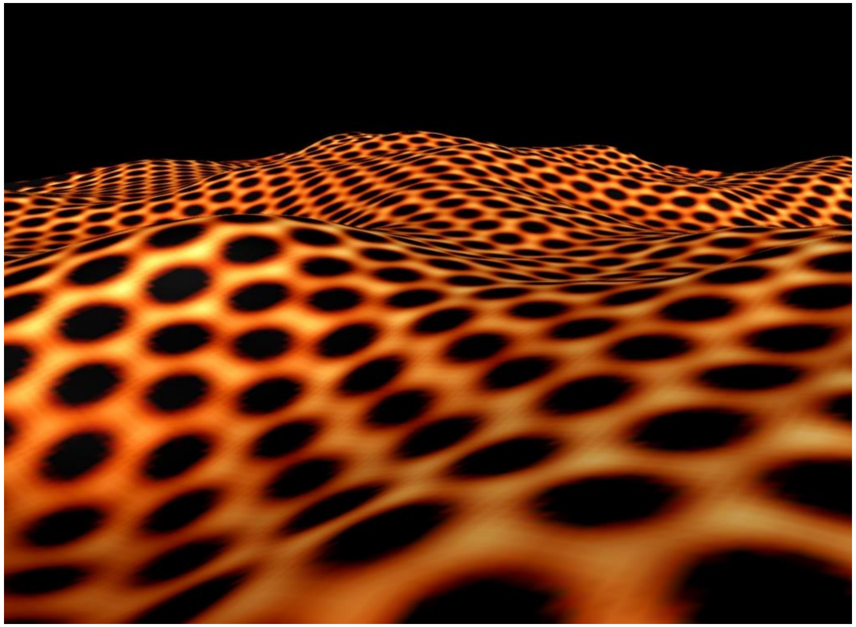
Graphene Curvature

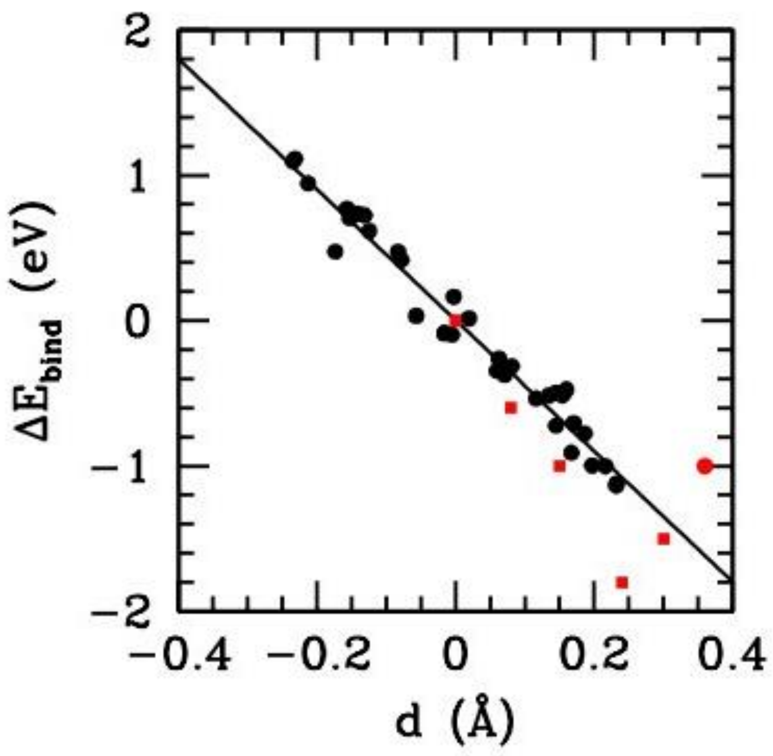
- Exploit graphene curvature for hydrogen storage at room temperature and pressure



V. Tozzini and V. Pellegrini: J. Phys. Chem. C **115**, 25523 (2011).

Graphene Curvature

- Exploit graphene curvature for hydrogen storage at room temperature and pressure
- The hydrogen binding energy on graphene is strongly dependent on local curvature and it is larger on convex parts



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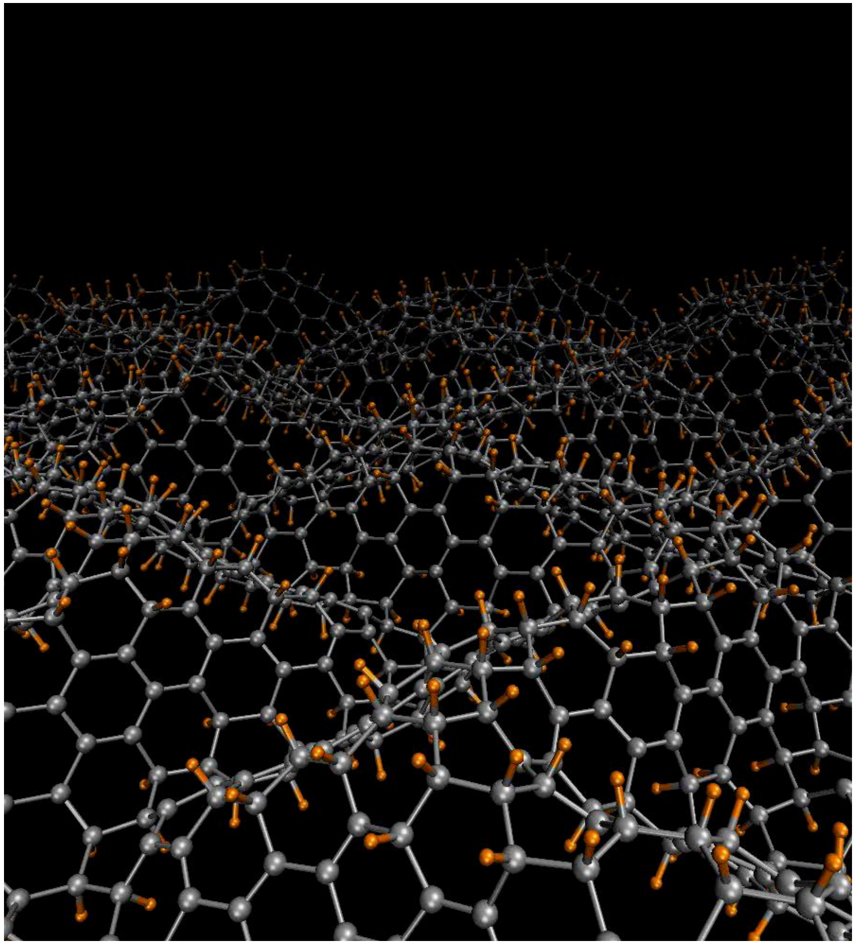
Graphene Curvature

- Exploit graphene curvature for hydrogen storage at room temperature and pressure
- The hydrogen binding energy on graphene is strongly dependent on local curvature and it is larger on convex parts
- Atomic hydrogen spontaneously sticks on convex parts; inverting curvature H is expelled



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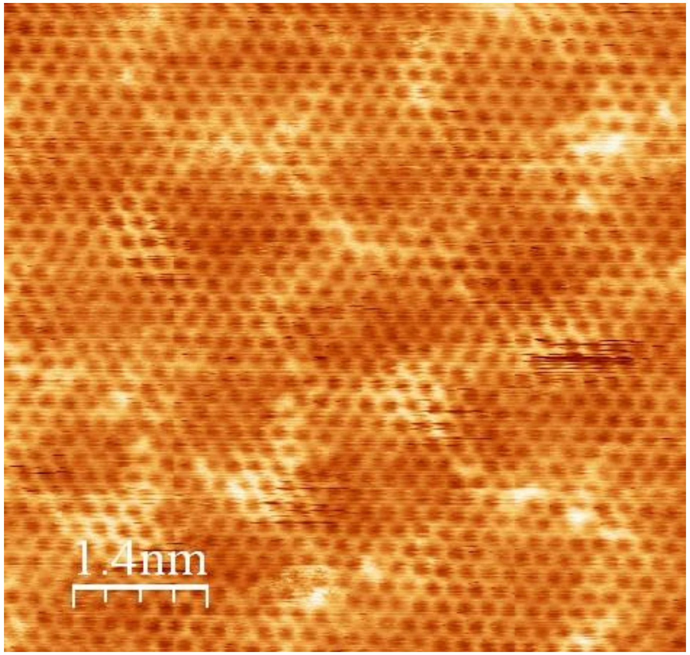
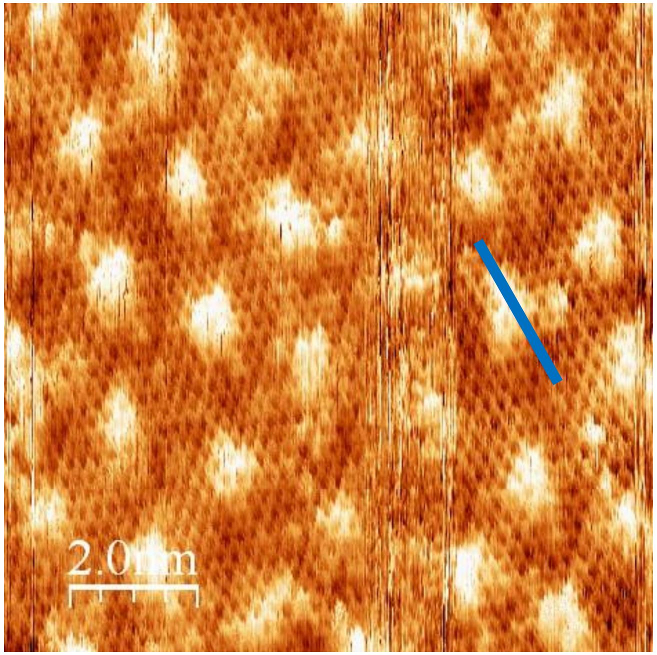
Hydrogen on graphene





Clean surface

Hydrogen clusters on corrugated graphene: preferential H binding on convex sites

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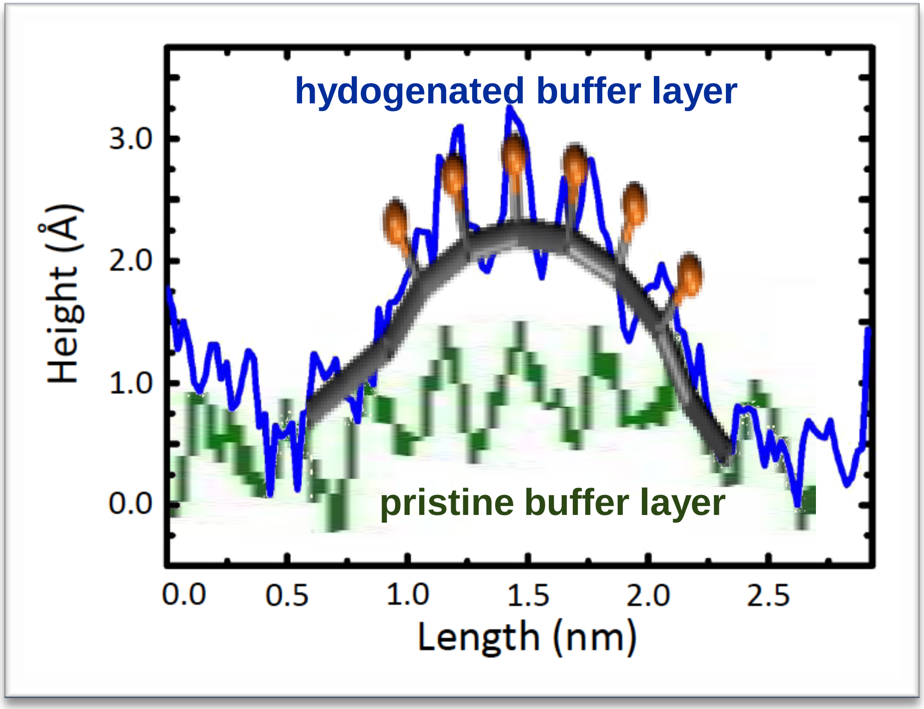
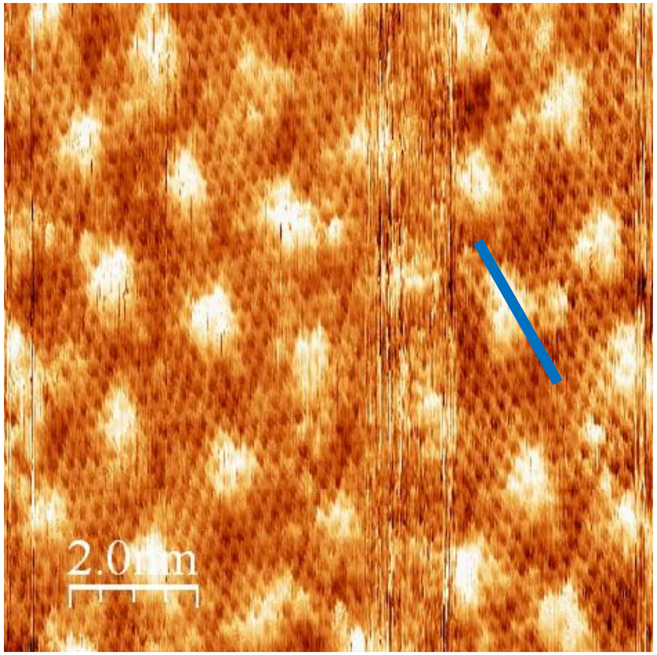
Hydrogen on graphene





Hydrogen clusters on corrugated graphene:
preferential H binding on convex sites

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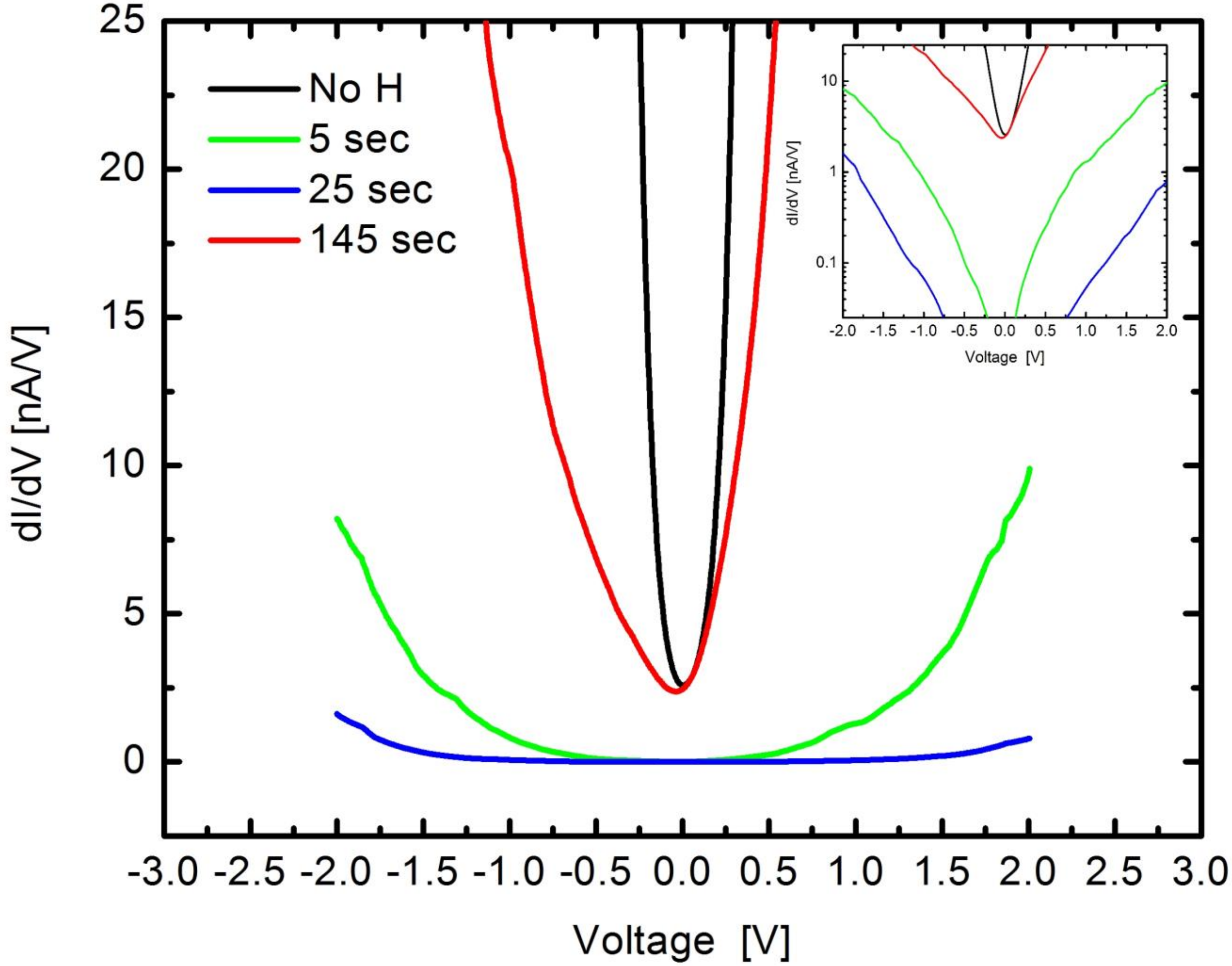
STS after hydrogenation



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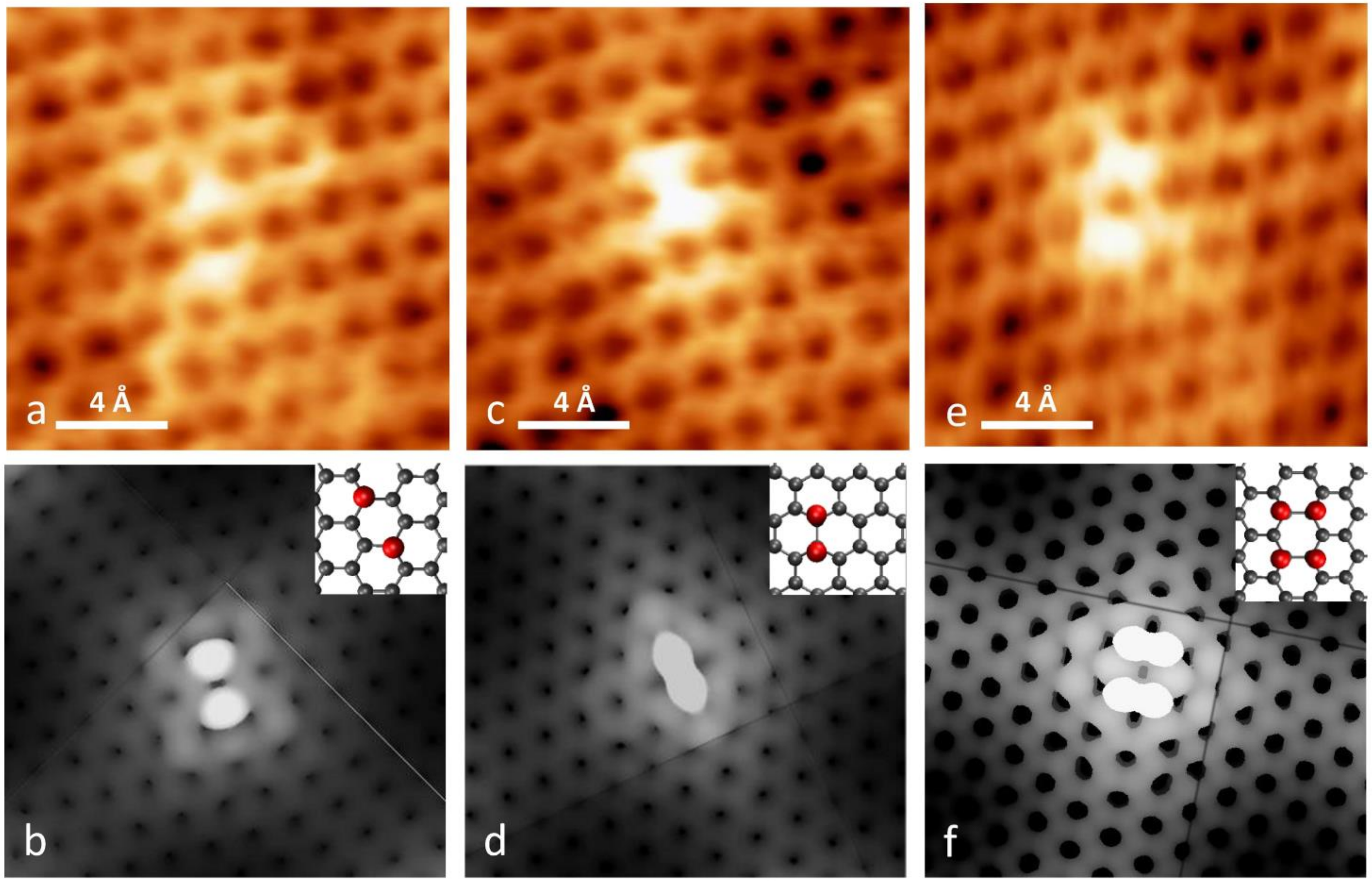
H-dimers and tetramers



Para-dimer | Ortho-dimer | Tetramer

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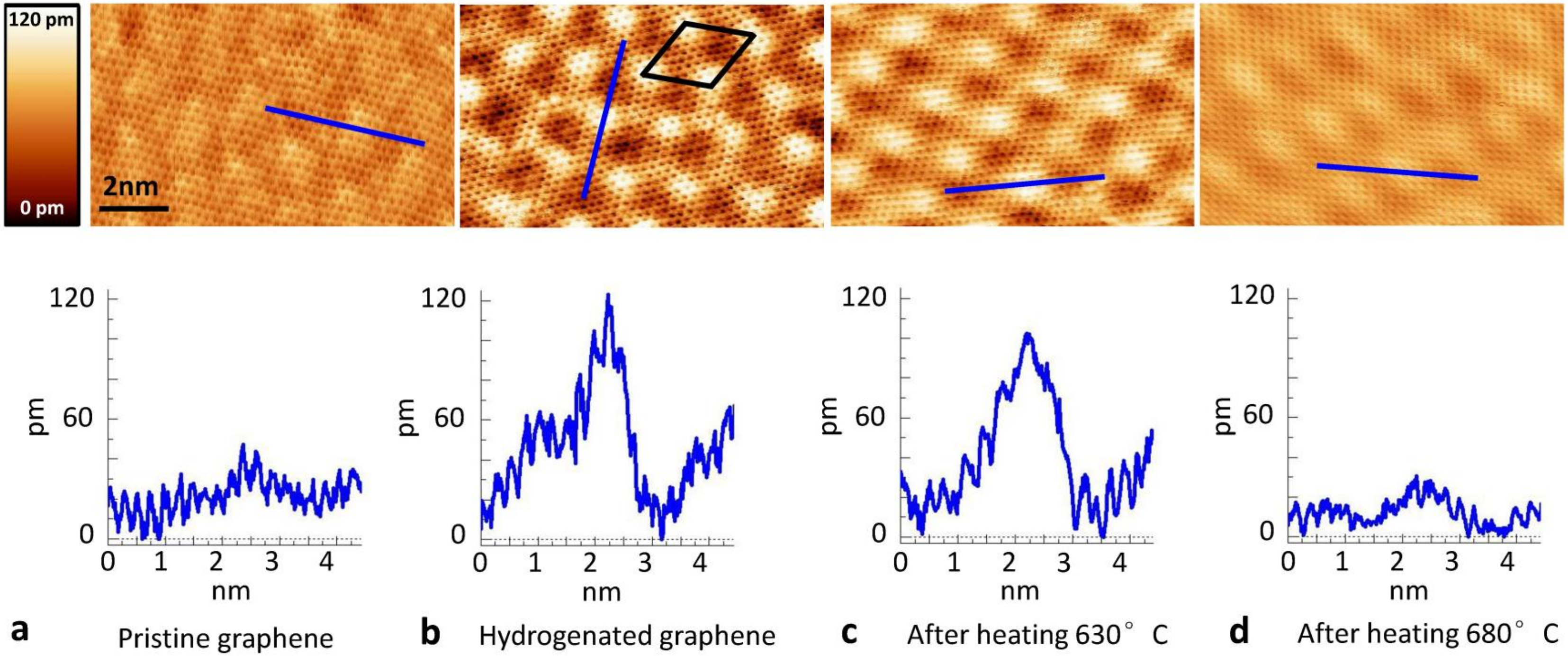
H adsorption and desorption



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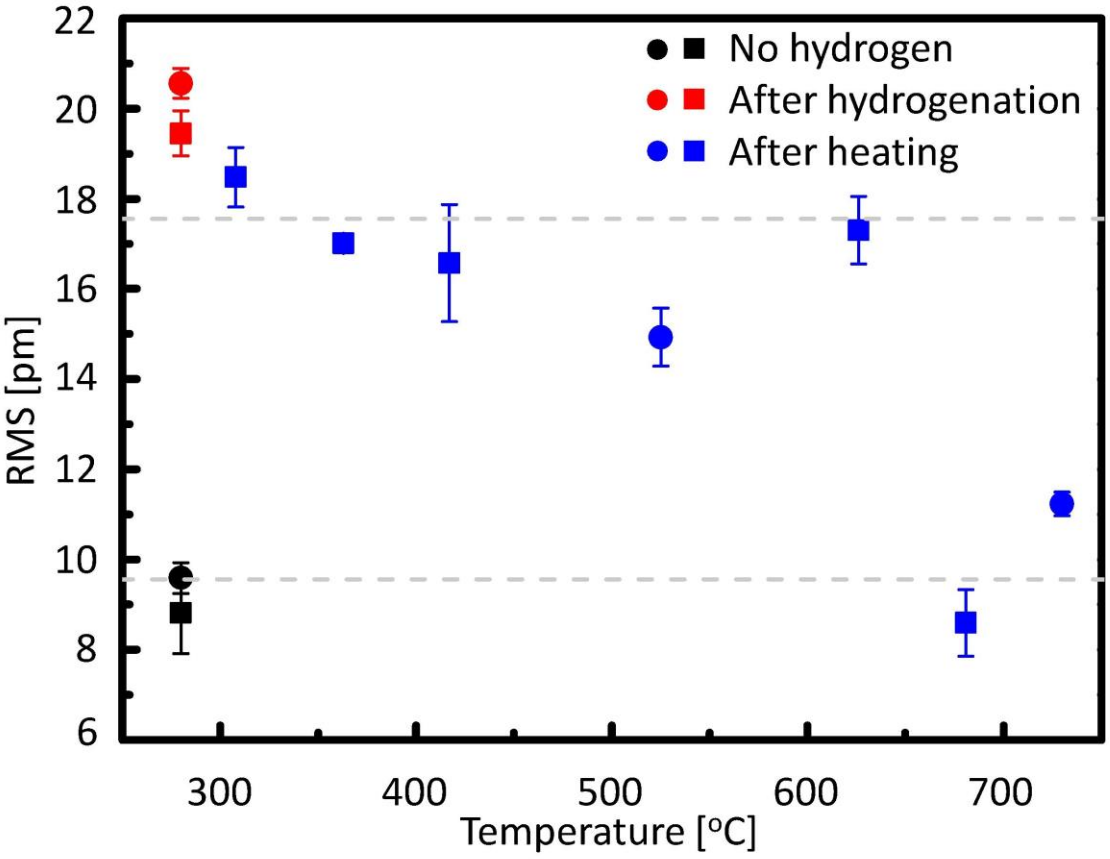
RMS roughness



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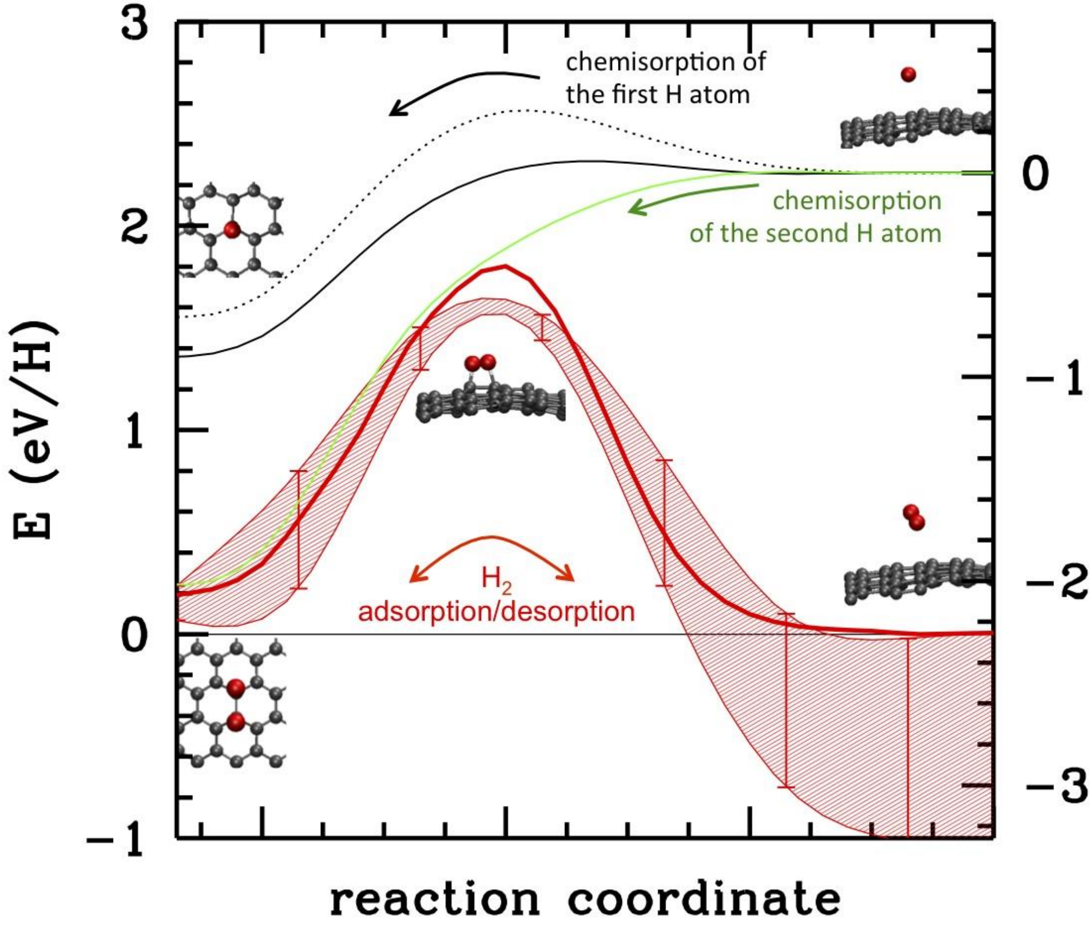
DFT calculations



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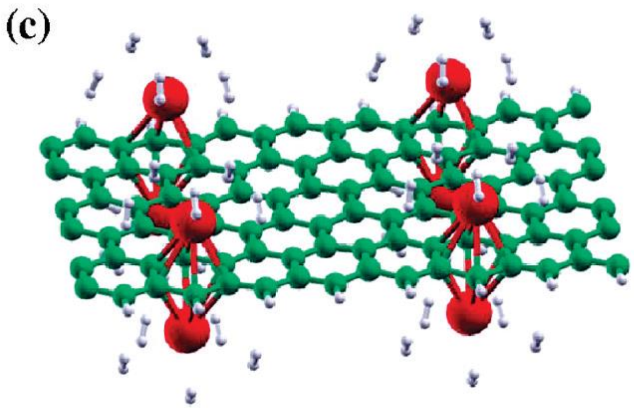
In collaboration with M. Takamura, S. Tanabe, H. Hibino

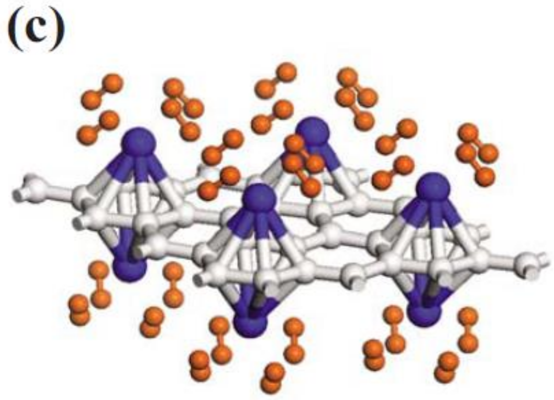
NTT Basic Research Laboratories, Atsugi, Japan

Functionalized Graphene

- Functionalized graphene has been predicted to adsorb up to 9 wt% of hydrogen
- Modify graphene with various chemical species, such as calcium or transition metals (Titanium)

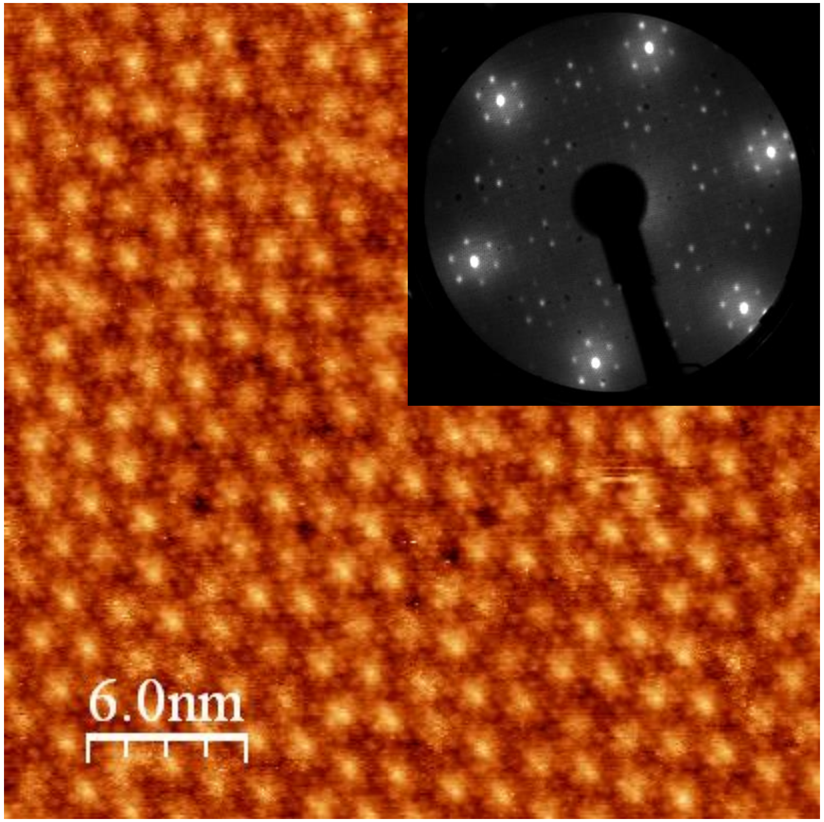


Lee et al., Nano Lett. 10 (2010) 793

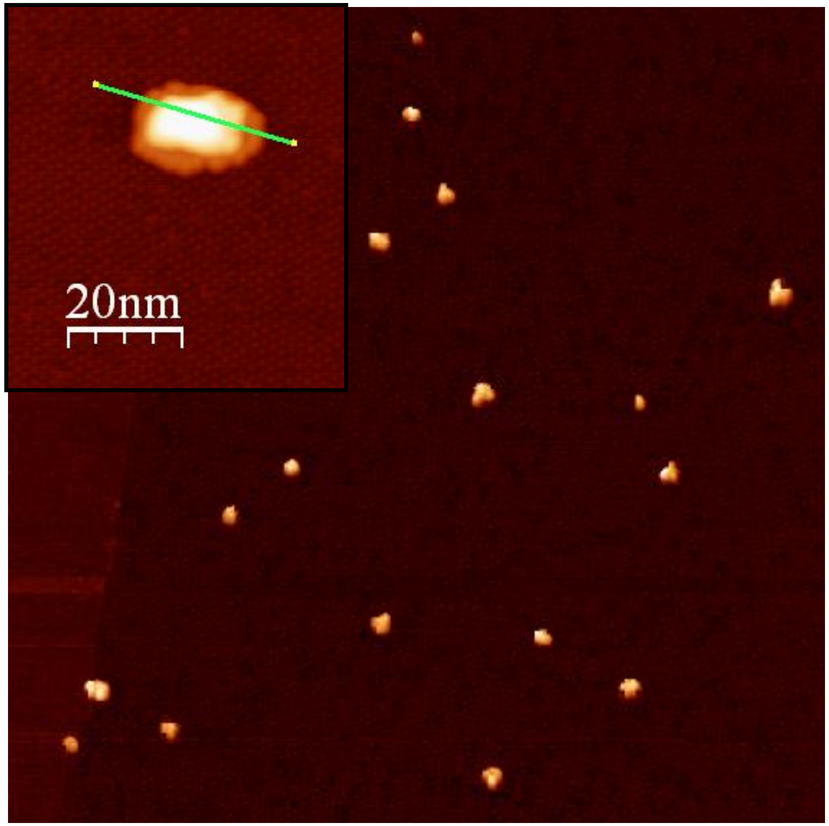


Durgen et al., PRB 77 (2007) 085405

Titanium on graphene



ML graphene on SiC(0001)
with reconstruction



After deposition of Ti at RT

T. Mashoff *et al.*

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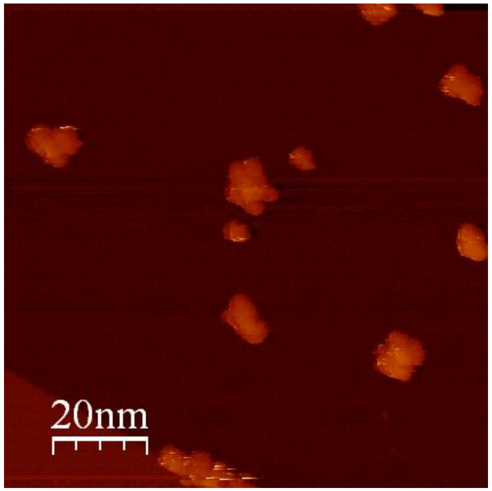
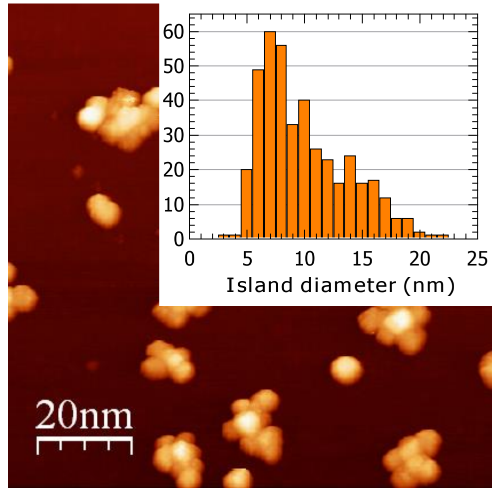
NEST

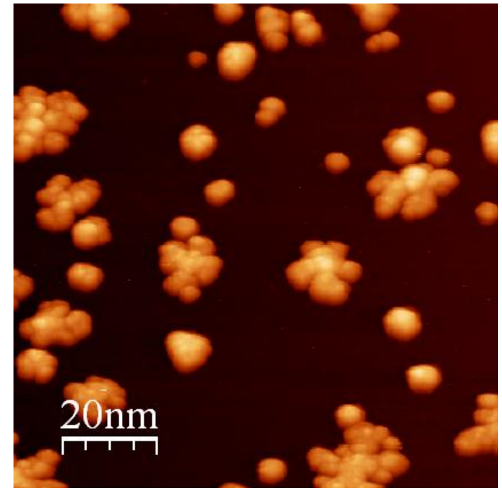
Titanium island growth



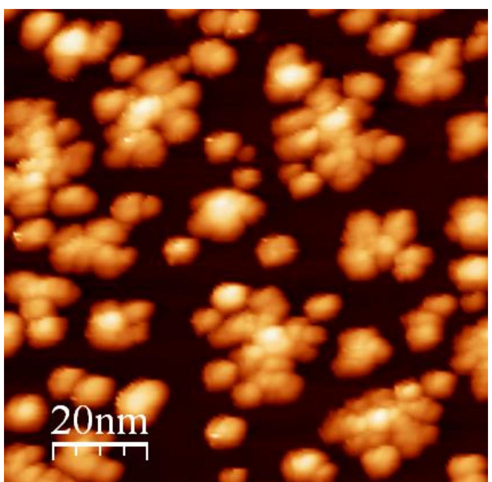
6% Coverage



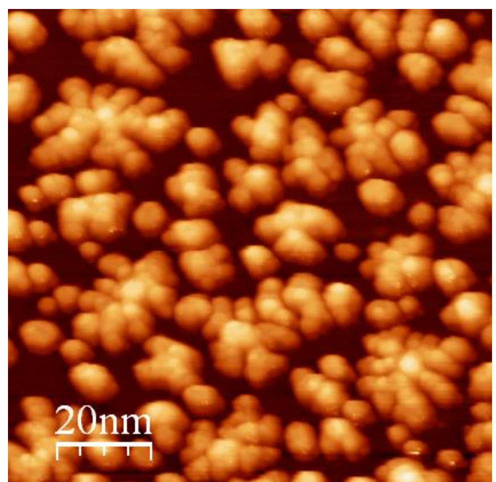
16% Coverage



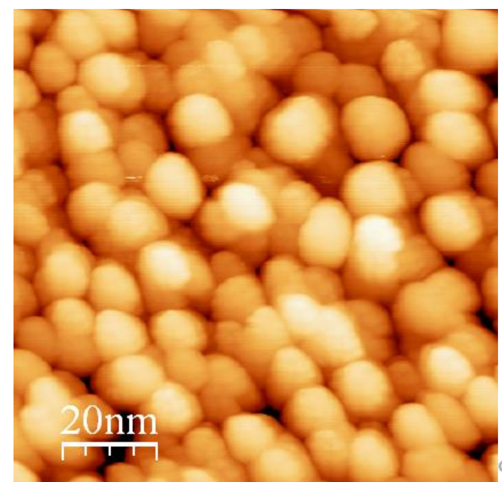
29% Coverage



53% Coverage



79% Coverage

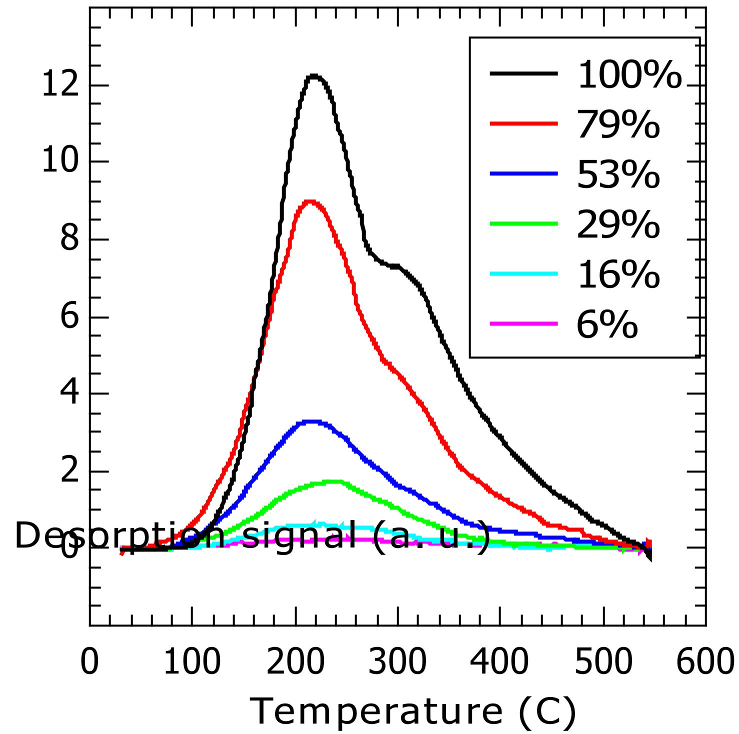


100% Coverage

Thermal desorption spectroscopy

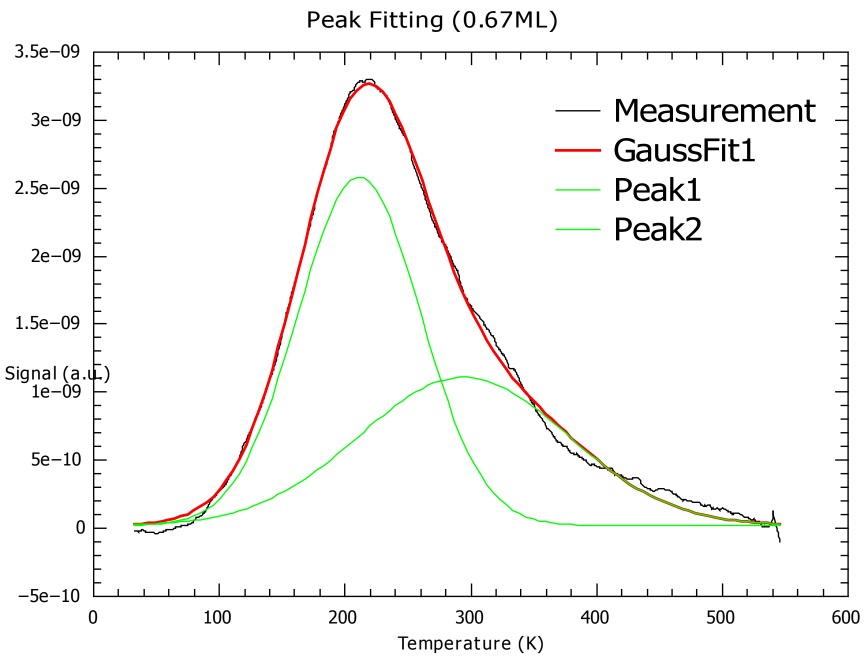
- Deposition of different amounts of Titanium
- Offering Hydrogen (D_2)
- ($1x10^{-7}$ mbar for 5 min)
- Heating sample with constant rate (10K/s) up to 550° C
- Measuring mass-sensitive desorption with a mass spectrometer

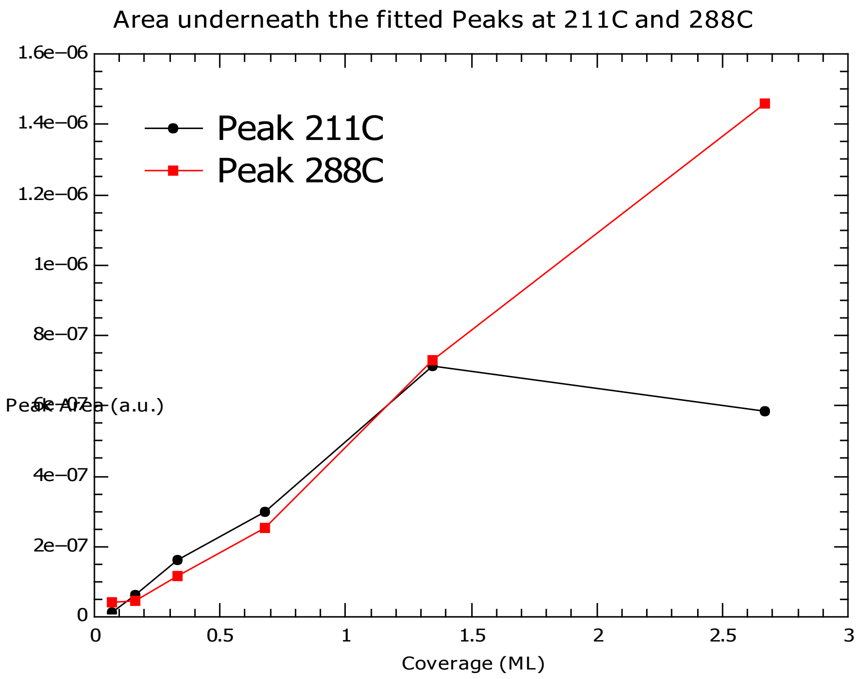
Spectra for different Ti-coverages



TDS peak intensity

Peak fitting





All curves could be fitted nicely
with 2 peaks at 211C and 288 C

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Conclusions

- Graphene is a promising material for hydrogen storage
- Curvature-dependent adsorption and desorption of hydrogen
  - reusable hydrogen storage devices that do not depend on temperature or pressure changes.
- Graphene functionalized by Ti:
  - Stability of hydrogen binding at room temperature
  - Hydrogen desorbes at moderate temperatures